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Average Reaction Rate Equation average reaction rate $= -\frac{\Delta[\text{reactant}]}{\Delta t}$ $\Delta[\text{reactant}]$ represents the change in concentration of a reactant. Δt represents the change in time. The average reaction rate for the consumption of a reactant is the negative change in the concentration of the reactant divided by the elapsed time. EXAMPLE Problem 16.1

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Chapter 16: Reaction Rates

CHM 116. 16. Chemical kinetics focuses on reaction rate and the factors that affect it

1. Concentration of reactants
2. Their physical state
3. Temperature of the reaction
4. Use of a catalyst

Rate Collision Frequency Concentration Under a given set of conditions, each reaction has its own rate Concentration affects the rate by influencing the frequency of collisions between reactant molecules

o Reaction rate is proportional to the number of collisions, ...

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248 Study Guide for An Introduction to Chemistry Exercise 16.2 – Equilibrium Constant Calculation: Ethanol, C_2H_5OH , can be made from the reaction of ethylene gas, C_2H_4 , and water vapor. A mixture of $C_2H_4(g)$ and $H_2O(g)$ is allowed to come to equilibrium in a container at $110\text{ }^\circ\text{C}$, and the partial pressures of

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the gases are found

Chapter 16 - The Process of Chemical Reactions

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A chemistry book lists the rate constants for two different reactions. The first rate constant is 0.12 1/s and the second is $34.5 \text{ 1/(M}\cdot\text{s)}$. Even though the rate equations for these reactions are not given, an astute student says that the overall order of these two chemical reactions must be different from each other.

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03 Kinetics Study Guide - Multiple Choice - Page 1 of 21 . Multiple
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CHM 152 Reaction Rates - Reading Guide Rates of Chemical Reactions Rate = change in concentration of reactant divided by change in time The general unit for rate is moles As time progresses, the concentration of the reactants _____ (decreases/increases). As time progresses, the concentration of the products _____ (decreases/ increases).

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