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Class Date Problem Solving continued
Sample Problem 2 In an experiment to

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identify an unknown gas, it is found that 1.82 L of the gas has a mass of 5.430 g. What is the density of the gas in g/L?
Solution ANALYZE What is given in the problem? the measured mass and volume of the gas

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CONVERTING NUMBER OF ATOMS OF AN
ELEMENT TO MASS Sample Problem 4
uses the progression of steps 1→2→3 to
convert from the mass of an element to
the number of atoms. In order to
calculate the mass from a given number
of atoms, these steps will be reversed.

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Chemists have many methods for
determining the quantity of a substance
present in a solution or other mixture.
One common method is titration, in
which a solution of known concentration

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reacts with a sample containing the
substance of unknown quantity.

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Workbook 1 Conversions Conversions

One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed

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into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how much, how big, how hot, how many, how hard, and how long did it take.

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PROBLEMS INVOLVING ATOMS AND
ELEMENTS Sample Problem 1 A chemist
has a jar containing 388.2 g of iron
filings. How many moles of iron does the
jar contain? Solution ANALYZE What is
given in the problem? mass of iron in
grams

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continued - PC\|MAC

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Problem Solving continued Sample
Problem 1 A HCl solution has a
concentration of 0.0050 M. Calculate
[OH⁻] and [H⁺] for this solution. HCl is
a strong acid, so assume it is 100% ion-
ized. Solution ANALYZE What is given in

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the problem? the molarity of the HCl
solution, and the fact

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Problem 1 Ammonia is made industrially
by reacting nitrogen and hydrogen

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Answer

under pressure, at high temperature, and in the presence of a catalyst. The equation is $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$. If 4.0 mol of H_2 react, how many moles of NH_3 will be produced?

Skills Worksheet Problem Solving

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Stoichiometry. So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of moles and that they relate moles of a substance to mass by using the molar mass.

Skills Worksheet Problem Solving - Mole Cafe

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One of the aims of chemistry is to describe changes—to tell what changed, how it changed, and what it changed into. Another aim of chemistry is to look at matter and its changes and to ask questions such as how much, how big, how hot, how many, how hard, and how

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Workbook 201 Concentration of
Solutions Sample Problem 3 Determine
the molal concentration of a solution
containing 81.3 g of ethylene glycol,

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HOCH₂CH₂OH, dissolved in 166 g of water. Solution ANALYZE What is given in the problem? the mass of ethylene glycol dissolved, and What are you asked to find? Items Data

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answer reasonable? Yes; 2 g of boron is
about $1/5$ of the molar mass of boron.
Therefore, 2.00 g boron will contain
about $1/5$ of an Avogadro's constant of
atoms. Practice 1.

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the balanced chemical equation Mass of
substance A Amount in mol of substance
A Amount in mol of substance B Convert
using the molar mass of A 1

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